

5 Microphysics of Cold Clouds

If a cloud extends above the *freezing level* (0°C level) it is called a *cold cloud*.

5 Microphysics of Cold Clouds

If a cloud extends above the *freezing level* (0°C level) it is called a *cold cloud*.

Even though the temperature may be below 0°C , water droplets can still exist in clouds, in which case they are referred to as *supercooled droplets*.

5 Microphysics of Cold Clouds

If a cloud extends above the *freezing level* (0°C level) it is called a *cold cloud*.

Even though the temperature may be below 0°C , water droplets can still exist in clouds, in which case they are referred to as *supercooled droplets*.

Cold clouds may also contain ice particles. If a cold cloud contains both ice particles and supercooled droplets it is said to be a *mixed cloud*.

5 Microphysics of Cold Clouds

If a cloud extends above the *freezing level* (0°C level) it is called a *cold cloud*.

Even though the temperature may be below 0°C , water droplets can still exist in clouds, in which case they are referred to as *supercooled droplets*.

Cold clouds may also contain ice particles. If a cold cloud contains both ice particles and supercooled droplets it is said to be a *mixed cloud*.

In this section we are concerned with the origins and concentrations of ice particles in clouds, the ways ice particles grow, and the formation of precipitation in cold clouds.

Nucleation of Ice Particles

A supercooled droplet is in an **unstable state**. For freezing to occur, enough water molecules must come together within the droplet to form an embryo of ice large enough to survive and grow.

Nucleation of Ice Particles

A supercooled droplet is in an **unstable state**. For freezing to occur, enough water molecules must come together within the droplet to form an embryo of ice large enough to survive and grow.

If a water droplet contains no foreign particles it can freeze only by *homogeneous nucleation*.

Nucleation of Ice Particles

A supercooled droplet is in an **unstable state**. For freezing to occur, enough water molecules must come together within the droplet to form an embryo of ice large enough to survive and grow.

If a water droplet contains no foreign particles it can freeze only by *homogeneous nucleation*.

Homogeneous nucleation occurs at about -41°C for droplets about $1\text{ }\mu\text{m}$ in diameter, and at about -35°C for drops $100\text{ }\mu\text{m}$ in diameter.

Nucleation of Ice Particles

A supercooled droplet is in an **unstable state**. For freezing to occur, enough water molecules must come together within the droplet to form an embryo of ice large enough to survive and grow.

If a water droplet contains no foreign particles it can freeze only by *homogeneous nucleation*.

Homogeneous nucleation occurs at about -41°C for droplets about $1\text{ }\mu\text{m}$ in diameter, and at about -35°C for drops $100\text{ }\mu\text{m}$ in diameter.

Hence, in the atmosphere, homogeneous nucleation of freezing occurs only in *high clouds*.

Heterogeneous Nucleation

If a droplet contains a rather special type of particle, called a *freezing nucleus*, it may freeze by a process known as *heterogeneous nucleation*:

Heterogeneous Nucleation

If a droplet contains a rather special type of particle, called a *freezing nucleus*, it may freeze by a process known as *heterogeneous nucleation*:

Water molecules in the droplet collect onto the surface of the particle to form an ice-like structure that may increase in size and cause the droplet to freeze.

Heterogeneous Nucleation

If a droplet contains a rather special type of particle, called a *freezing nucleus*, it may freeze by a process known as *heterogeneous nucleation*:

Water molecules in the droplet collect onto the surface of the particle to form an ice-like structure that may increase in size and cause the droplet to freeze.

Since the formation of the ice structure is aided by the freezing nucleus, heterogeneous nucleation can occur at **much higher temperatures** than homogeneous nucleation.

Heterogeneous Nucleation

If a droplet contains a rather special type of particle, called a *freezing nucleus*, it may freeze by a process known as *heterogeneous nucleation*:

Water molecules in the droplet collect onto the surface of the particle to form an ice-like structure that may increase in size and cause the droplet to freeze.

Since the formation of the ice structure is aided by the freezing nucleus, heterogeneous nucleation can occur at **much higher temperatures** than homogeneous nucleation.

Certain particles in the air also serve as centers upon which ice can form **directly from the vapour phase**. These particles are referred to as *deposition nuclei*.

Heterogeneous Nucleation

If a droplet contains a rather special type of particle, called a *freezing nucleus*, it may freeze by a process known as *heterogeneous nucleation*:

Water molecules in the droplet collect onto the surface of the particle to form an ice-like structure that may increase in size and cause the droplet to freeze.

Since the formation of the ice structure is aided by the freezing nucleus, heterogeneous nucleation can occur at **much higher temperatures** than homogeneous nucleation.

Certain particles in the air also serve as centers upon which ice can form **directly from the vapour phase**. These particles are referred to as *deposition nuclei*.

Ice can form by deposition if the air is supersaturated with respect to ice and the temperature is low enough.

Ice Nuclei

Particles with molecular spacings and crystallographic arrangements similar to those of ice (which has a hexagonal structure) tend to be effective as ice nuclei.

Ice Nuclei

Particles with molecular spacings and crystallographic arrangements similar to those of ice (which has a hexagonal structure) tend to be effective as ice nuclei.

Most effective ice nuclei are virtually **insoluble** in water.

Ice Nuclei

Particles with molecular spacings and crystallographic arrangements similar to those of ice (which has a hexagonal structure) tend to be effective as ice nuclei.

Most effective ice nuclei are virtually **insoluble** in water.

Some inorganic soil particles (mainly clays) can nucleate ice at fairly high temperatures (i.e., above -15°C), and play an important role in nucleating ice in clouds.

Measuring Ice Nuclei Concentration

Several techniques have been used for **measuring the concentrations** of particles in the air that are active as ice nuclei at a given temperature.

Measuring Ice Nuclei Concentration

Several techniques have been used for **measuring the concentrations** of particles in the air that are active as ice nuclei at a given temperature.

A common method is to draw a known volume of air into a container and to cool it until a cloud is formed. The number of ice crystals forming at a particular temperature is then measured.

Measuring Ice Nuclei Concentration

Several techniques have been used for **measuring the concentrations** of particles in the air that are active as ice nuclei at a given temperature.

A common method is to draw a known volume of air into a container and to cool it until a cloud is formed. The number of ice crystals forming at a particular temperature is then measured.

In *expansion chambers* cooling is produced by compressing the air and then suddenly expanding it.

Measuring Ice Nuclei Concentration

Several techniques have been used for **measuring the concentrations** of particles in the air that are active as ice nuclei at a given temperature.

A common method is to draw a known volume of air into a container and to cool it until a cloud is formed. The number of ice crystals forming at a particular temperature is then measured.

In *expansion chambers* cooling is produced by compressing the air and then suddenly expanding it.

In *mixing chambers* cooling is produced by refrigeration.

Measuring Ice Nuclei Concentration

Several techniques have been used for **measuring the concentrations** of particles in the air that are active as ice nuclei at a given temperature.

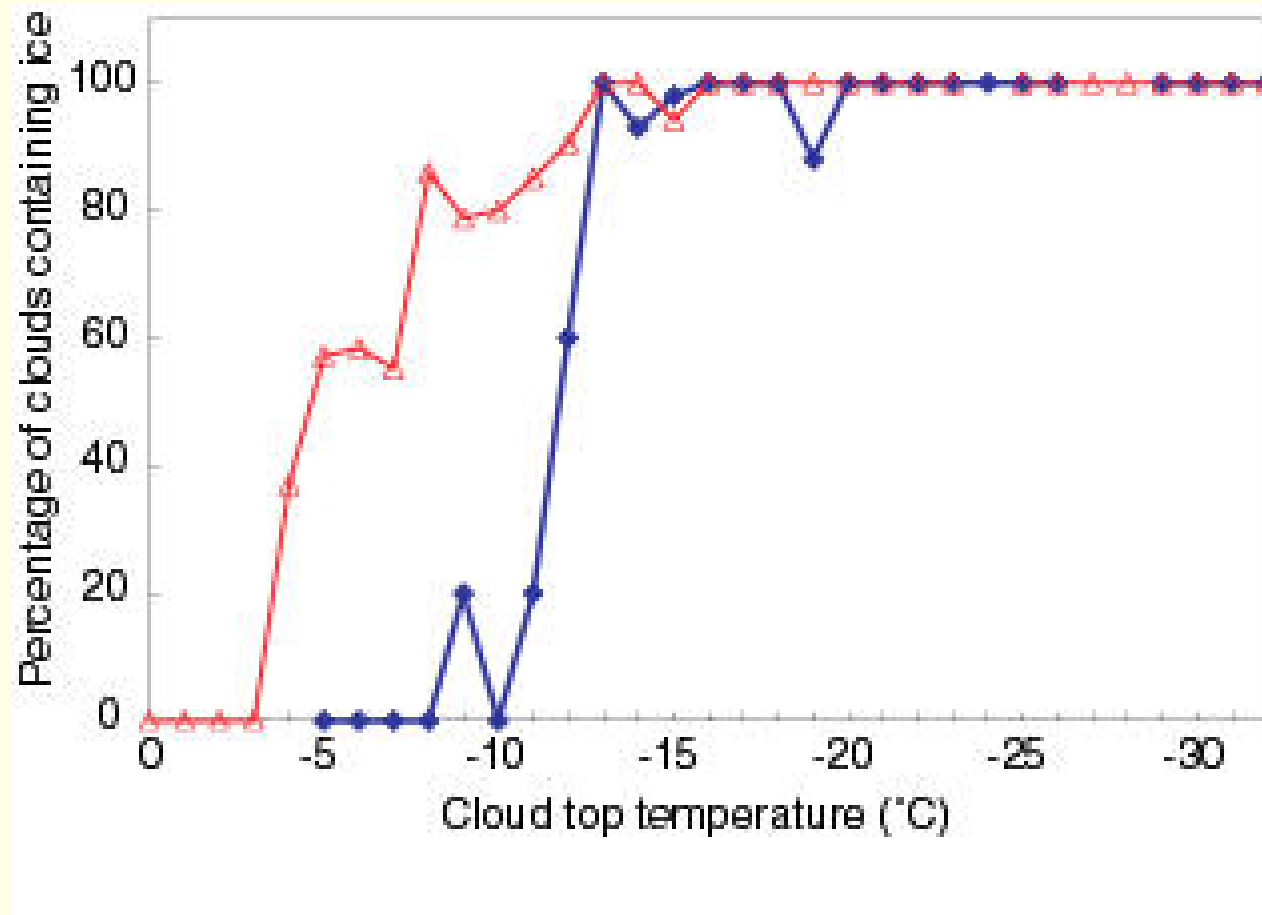
A common method is to draw a known volume of air into a container and to cool it until a cloud is formed. The number of ice crystals forming at a particular temperature is then measured.

In *expansion chambers* cooling is produced by compressing the air and then suddenly expanding it.

In *mixing chambers* cooling is produced by refrigeration.

In *diffusion chambers* temperature, supersaturation and pressure can be controlled independently.

Concentrations of Ice Particles



Percentage of clouds containing ice particle concentrations greater than about 1 per liter as a function of cloud top temperature. **Blue curve:** Continental cumuliform clouds. **Red curve:** Clean marine cumuliform clouds and clean arctic stratiform clouds.

The probability of ice particles being present in a cloud increases as the temperature decreases below 0°C , as illustrated in the Figure above.

The probability of ice particles being present in a cloud increases as the temperature decreases below 0°C , as illustrated in the Figure above.

The results indicate that the probability of ice being present is 100% for cloud top temperature below about -13°C .

The probability of ice particles being present in a cloud increases as the temperature decreases below 0°C , as illustrated in the Figure above.

The results indicate that the probability of ice being present is 100% for cloud top temperature below about -13°C .

At higher temperatures the probability of ice being present falls off sharply, but it is greater if the cloud contains drizzle or raindrops.

The probability of ice particles being present in a cloud **increases as the temperature decreases** below 0°C , as illustrated in the Figure above.

The results indicate that the probability of ice being present is 100% for cloud top temperature below about -13°C .

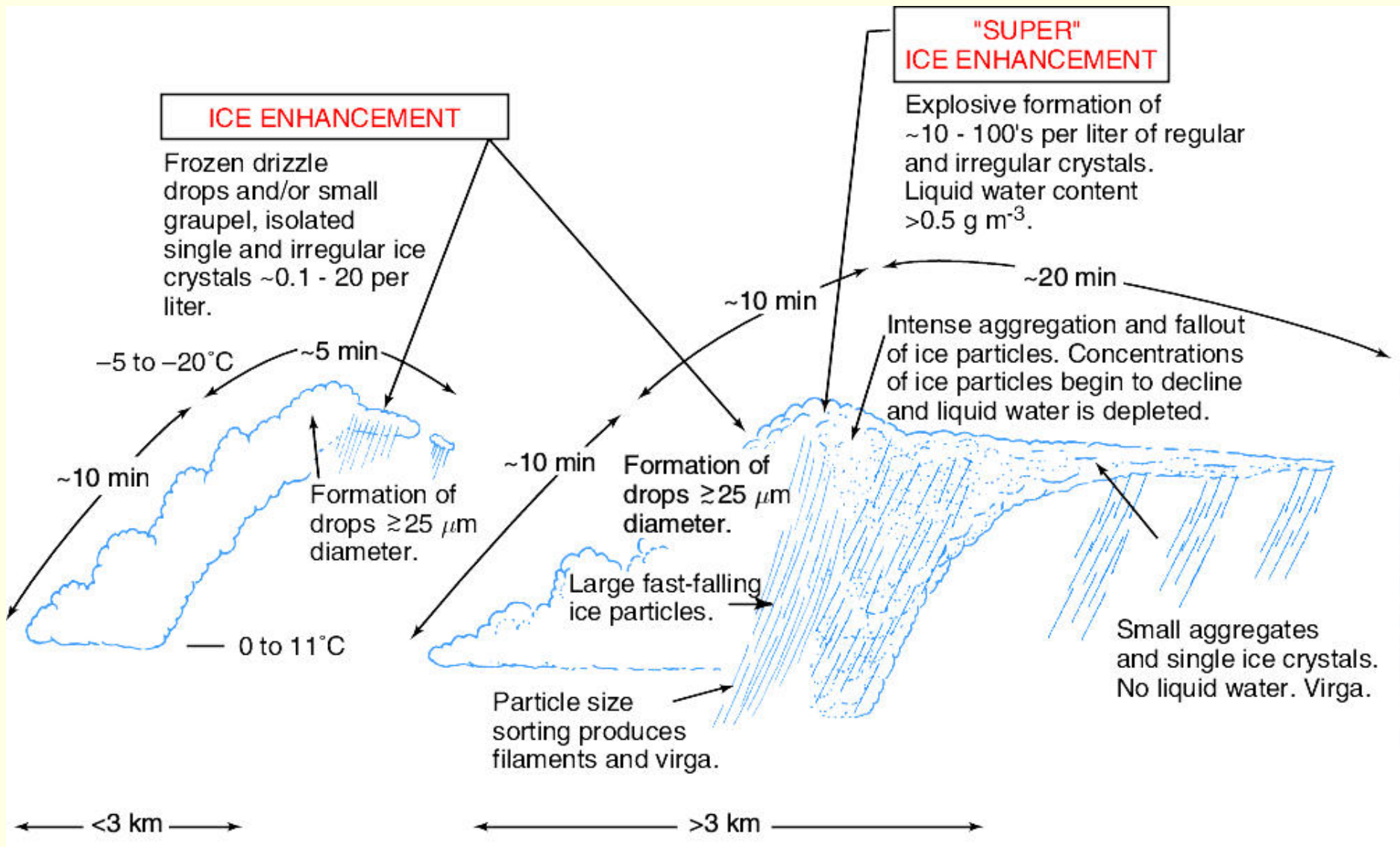
At higher temperatures the probability of ice being present falls off sharply, but it is greater if the cloud contains drizzle or raindrops.

Clouds with top temperatures between about 0 and -8°C generally contain copious supercooled droplets. It is in clouds such as these that aircraft are most likely to encounter **severe icing** conditions, since supercooled droplets freeze when they collide with an aircraft.

Growth of Ice Particles in Clouds

We consider several methods of growth:

- (a) Growth from the vapour phase.
- (b) Growth by riming; hailstones.
- (c) Growth by aggregation.



Schematic of ice development in small cumuliform clouds.

(a) *Growth from the vapour phase.*

In a mixed cloud dominated by supercooled droplets, the air is close to saturated with respect to liquid water and is therefore *supersaturated with respect to ice*.

(a) *Growth from the vapour phase.*

In a mixed cloud dominated by supercooled droplets, the air is close to saturated with respect to liquid water and is therefore *supersaturated with respect to ice*.

For example, air saturated with respect to liquid water at -10°C is supersaturated with respect to ice by 10%.

This value is much higher than the supersaturation of cloudy air with respect to liquid water, which rarely exceed 1%.

(a) *Growth from the vapour phase.*

In a mixed cloud dominated by supercooled droplets, the air is close to saturated with respect to liquid water and is therefore *supersaturated with respect to ice*.

For example, air saturated with respect to liquid water at -10°C is supersaturated with respect to ice by 10%.

This value is much higher than the supersaturation of cloudy air with respect to liquid water, which rarely exceed 1%.

Consequently, in mixed clouds dominated by supercooled water droplets, in which the cloudy air is close to water saturation, *ice particles will grow from the vapour phase much more rapidly than droplets*.

(a) *Growth from the vapour phase.*

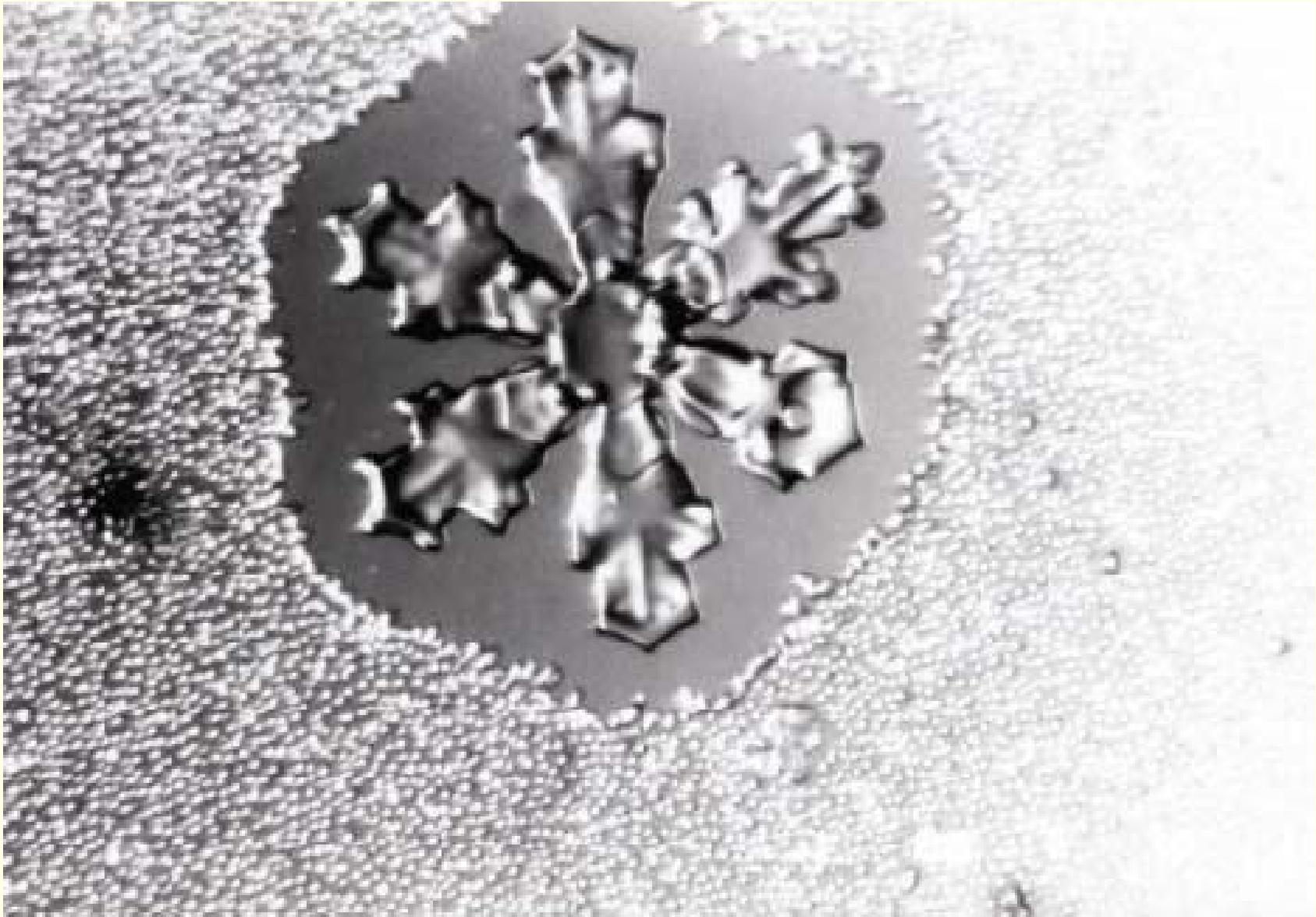
In a mixed cloud dominated by supercooled droplets, the air is close to saturated with respect to liquid water and is therefore *supersaturated with respect to ice*.

For example, air saturated with respect to liquid water at -10°C is supersaturated with respect to ice by 10%.

This value is much higher than the supersaturation of cloudy air with respect to liquid water, which rarely exceed 1%.

Consequently, in mixed clouds dominated by supercooled water droplets, in which the cloudy air is close to water saturation, *ice particles will grow from the vapour phase much more rapidly than droplets*.

In fact, if a growing ice particle lowers the vapour pressure in its vicinity below water saturation, adjacent droplets will evaporate (Figure below).

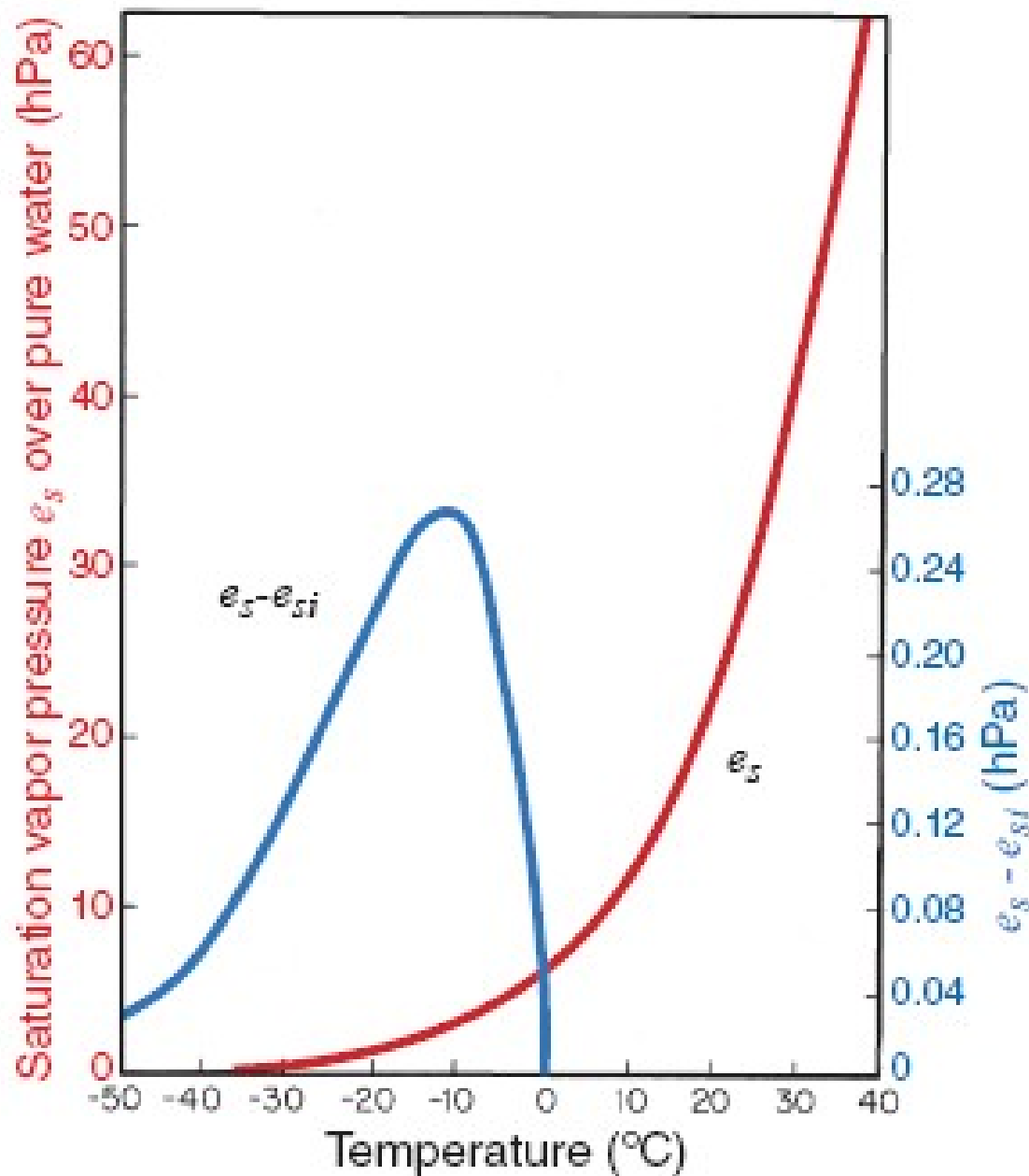


Laboratory demonstration of the growth of an ice crystal at the expense of surrounding supercooled water drops.



Cumulus turrets containing relatively large ice particles often have ill defined, fuzzy boundaries, as for the clouts in the background here.

Turrets containing only small droplets have well-defined, sharper boundaries, particularly if the cloud is growing (see cloud in foreground).



Saturation vapour pressure over water (Red) and the difference between the saturation pressures over water and over ice (Blue).

The lower equilibrium vapour pressure over ice than over water at the same temperature allows ice particles to migrate for greater distances than droplets into the non-saturated air surrounding a cloud before they evaporate.

The lower equilibrium vapour pressure over ice than over water at the same temperature allows ice particles to migrate for greater distances than droplets into the non-saturated air surrounding a cloud before they evaporate.

For this reason, ice particles that are large enough to fall out of a cloud can survive great distances before evaporating completely, even if the ambient air is sub-saturated with respect to ice.

The lower equilibrium vapour pressure over ice than over water at the same temperature allows ice particles to migrate for greater distances than droplets into the non-saturated air surrounding a cloud before they evaporate.

For this reason, ice particles that are large enough to fall out of a cloud can survive great distances before evaporating completely, even if the ambient air is sub-saturated with respect to ice.

The trails of ice crystals so produced are called *fallstreaks* or *virga*.

The lower equilibrium vapour pressure over ice than over water at the same temperature allows ice particles to migrate for greater distances than droplets into the non-saturated air surrounding a cloud before they evaporate.

For this reason, ice particles that are large enough to fall out of a cloud can survive great distances before evaporating completely, even if the ambient air is sub-saturated with respect to ice.

The trails of ice crystals so produced are called *fallstreaks* or *virga*.

Ice particles will grow in air that is sub-saturated with respect to water, provided that it is supersaturated with respect to ice.



Fallstreaks of ice crystals from cirrus clouds. The characteristic curved shape of the fallstreaks indicates that the wind speed was increasing (from left to right) with increasing altitude.

Shape and Form of Ice particles

The majority of ice particles in clouds are irregular in shape (sometimes referred to as “junk” ice).

Shape and Form of Ice particles

The majority of ice particles in clouds are irregular in shape (sometimes referred to as “junk” ice).

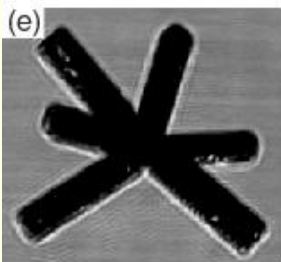
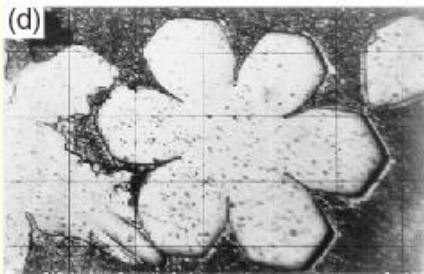
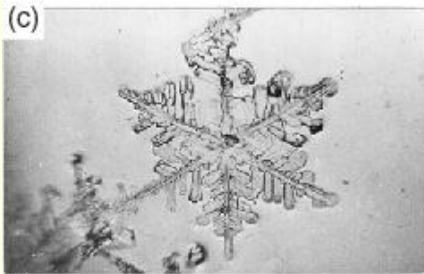
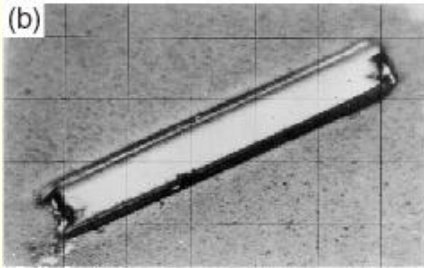
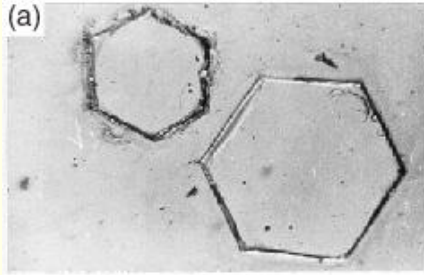
Laboratory studies show that under appropriate conditions ice crystals that grow from the vapour phase can assume a variety of regular shapes that are either platelike or column-like.

Shape and Form of Ice particles

The majority of ice particles in clouds are irregular in shape (sometimes referred to as “junk” ice).

Laboratory studies show that under appropriate conditions ice crystals that grow from the vapour phase can assume a variety of regular shapes that are either platelike or columnlike.

The simplest platelike crystals are plane *hexagonal plates*, and the simplest columnlike crystals are solid columns that are hexagonal in cross section.



Ice crystals grown
from the vapour phase:
(a) hexagonal plates,
(b) column,
(c) dendrite,
(d) sector plate,
(e) bullet rosette.

(b) *Growth by riming; hailstones.*

In a mixed cloud, ice particles can increase in mass by colliding with supercooled droplets which then freeze onto them.

(b) *Growth by riming; hailstones.*

In a mixed cloud, ice particles can increase in mass by colliding with supercooled droplets which then freeze onto them.

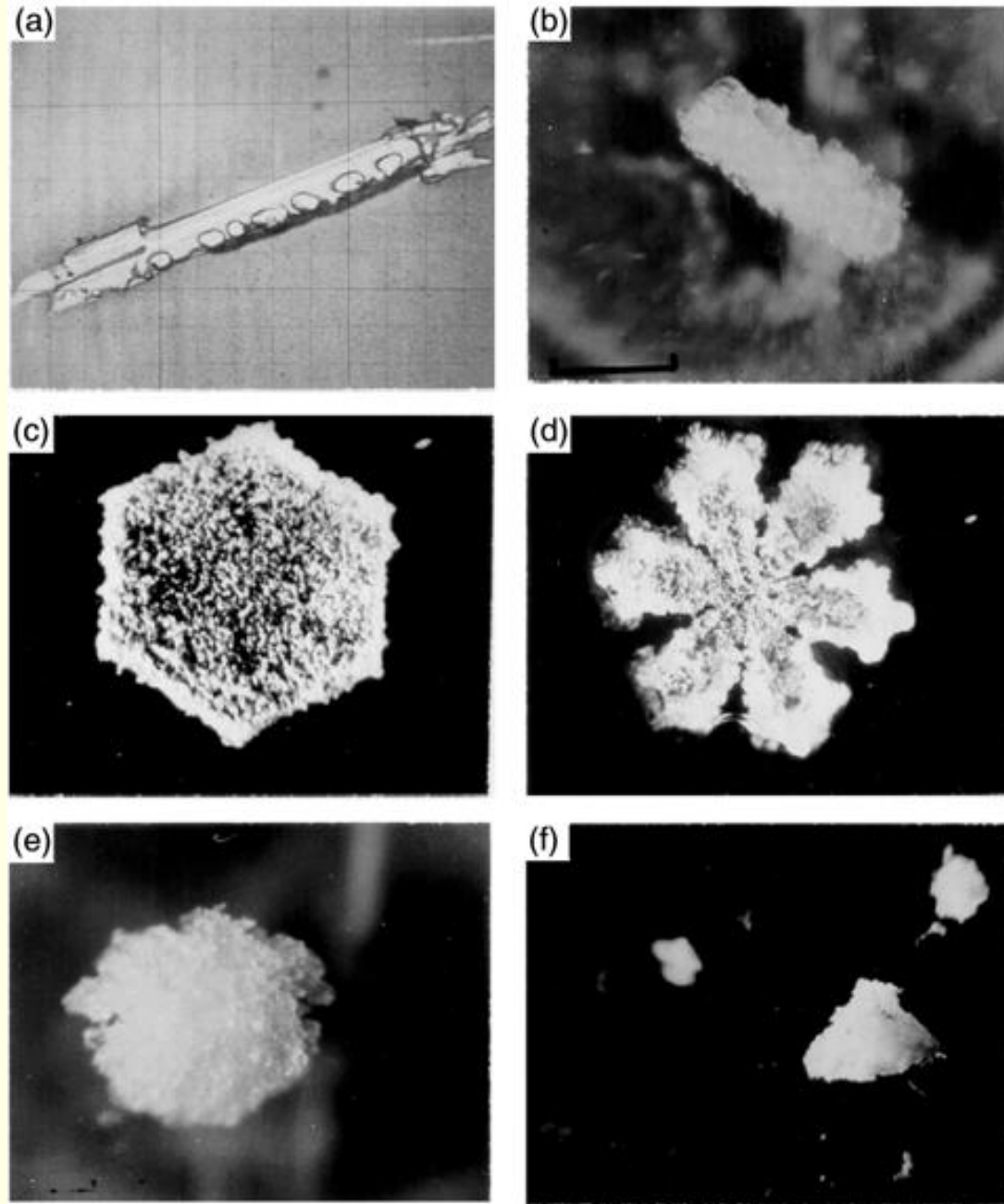
This process, referred to as *growth by riming*, leads to the formation of various rimed structures (figure follows).

(b) *Growth by riming; hailstones.*

In a mixed cloud, ice particles can increase in mass by colliding with supercooled droplets which then freeze onto them.

This process, referred to as *growth by riming*, leads to the formation of various rimed structures (figure follows).

When riming proceeds beyond a certain stage it becomes difficult to discern the original shape of the ice crystal. The rimed particle is then referred to as *graupel*.



(a) Lightly rimed needle; (b) rimed column; (c) rimed plate; (d) rimed stellar; (e) spherical graupel; (f) conical graupel.

Hailstones

Hailstones represent an extreme case of the growth of ice particles by riming. They form in vigorous convective clouds that have high liquid water contents.

Hailstones

Hailstones represent an extreme case of the growth of ice particles by riming. They form in vigorous convective clouds that have high liquid water contents.

The largest hailstone reported in the USA (Nebraska) was **13.8 cm in diameter** and weighed about 0.7 kg. However, hailstones about 1 cm in diameter are much more common.

Hailstones

Hailstones represent an extreme case of the growth of ice particles by riming. They form in vigorous convective clouds that have high liquid water contents.

The largest hailstone reported in the USA (Nebraska) was **13.8 cm in diameter** and weighed about 0.7 kg. However, hailstones about 1 cm in diameter are much more common.

If a thin section is cut from a hailstone and viewed in transmitted light, it is often seen to consist of alternate dark and light layers (Figure follows).

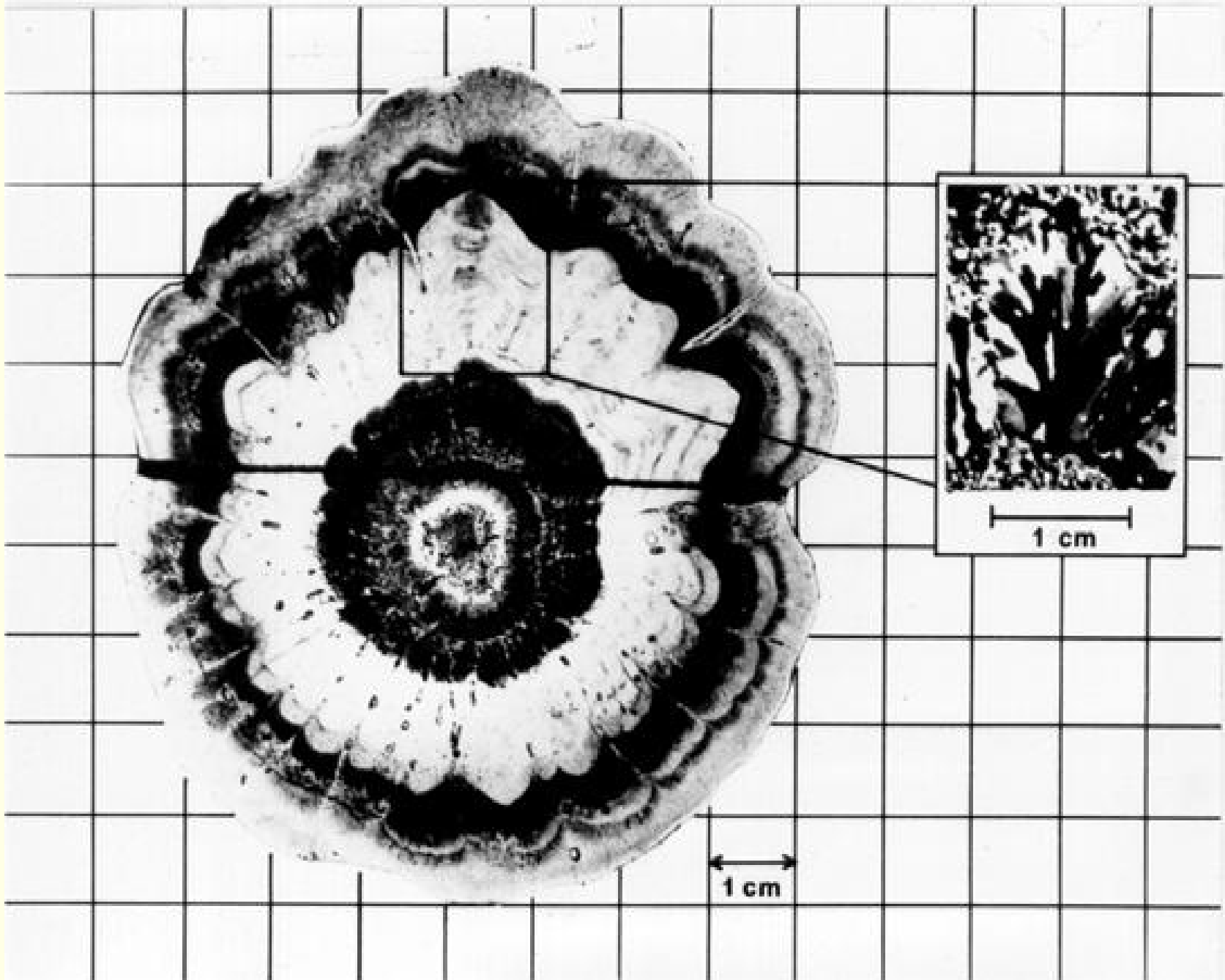
Hailstones

Hailstones represent an extreme case of the growth of ice particles by riming. They form in vigorous convective clouds that have high liquid water contents.

The largest hailstone reported in the USA (Nebraska) was **13.8 cm in diameter** and weighed about 0.7 kg. However, hailstones about 1 cm in diameter are much more common.

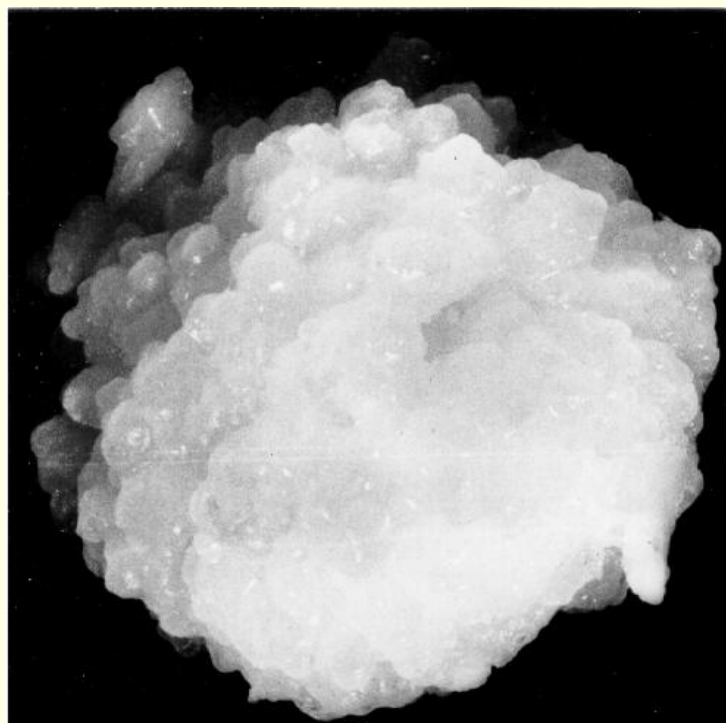
If a thin section is cut from a hailstone and viewed in transmitted light, it is often seen to consist of alternate dark and light layers (Figure follows).

The dark layers are opaque ice containing numerous small air bubbles, and the light layers are clear ice. Clear ice is more likely to form when the hailstone is growing wet.

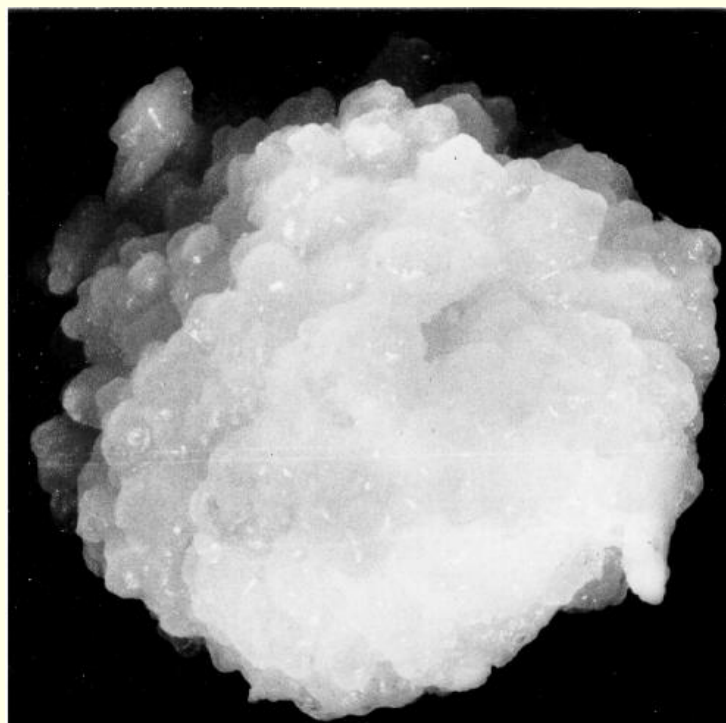


Thin section through the center of a hailstone.

Artificial hailstone (i.e., grown in the laboratory) showing a lobe structure. Growth was initially dry but tended toward wet growth as the stone grew.



Artificial hailstone (i.e., grown in the laboratory) showing a lobe structure. Growth was initially dry but tended toward wet growth as the stone grew.



The development of lobes in a hailstone may be due to the fact that small bumps on a hailstone will be areas of enhanced collection efficiencies for droplets.

(c) *Growth by aggregation.*

(c) *Growth by aggregation.*

The third mechanism by which ice particles grow in clouds is by *colliding* and *aggregating* with one another.

(c) *Growth by aggregation.*

The third mechanism by which ice particles grow in clouds is by *colliding* and *aggregating* with one another.

Ice particles can collide with each other provided their terminal fall speeds are different.

For example, graupel particles 1 and 4 mm in diameter have terminal fall speeds of about 1 and 2.5 m s^{-1} respectively.

(c) *Growth by aggregation.*

The third mechanism by which ice particles grow in clouds is by *colliding* and *aggregating* with one another.

Ice particles can collide with each other provided their terminal fall speeds are different.

For example, graupel particles 1 and 4 mm in diameter have terminal fall speeds of about 1 and 2.5 m s^{-1} respectively.

Consequently, the frequency of collisions of ice particles in clouds is greatly enhanced if some riming has taken place.

The second factor that influences growth by aggregation is whether or not two ice particles *adhere* when they collide.

The second factor that influences growth by aggregation is whether or not two ice particles *adhere* when they collide.

The probability of adhesion is determined primarily by:

1. The types of ice particles
2. The temperature

The second factor that influences growth by aggregation is whether or not two ice particles *adhere* when they collide.

The probability of adhesion is determined primarily by:

1. The types of ice particles
2. The temperature

The probability of two colliding crystals adhering increases with increasing temperature as the ice surfaces become more “sticky”.

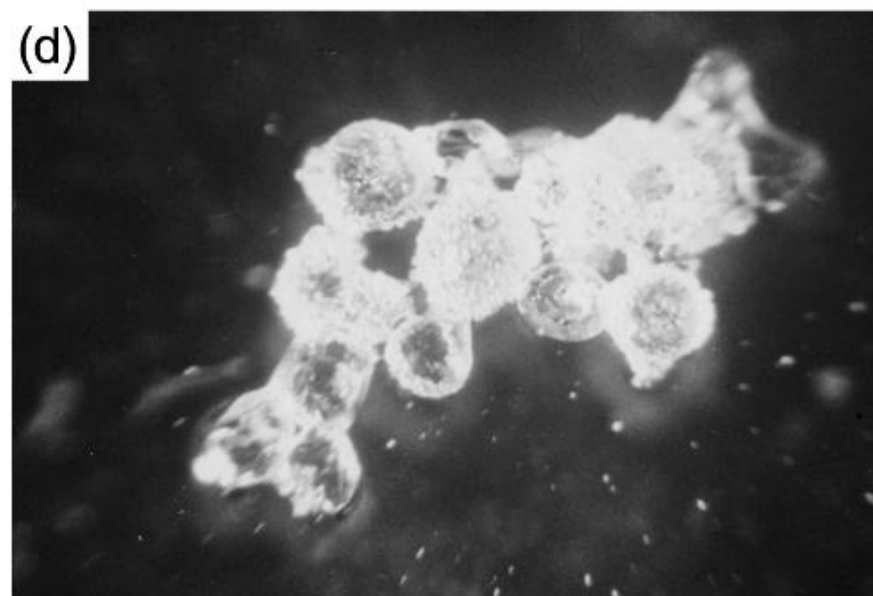
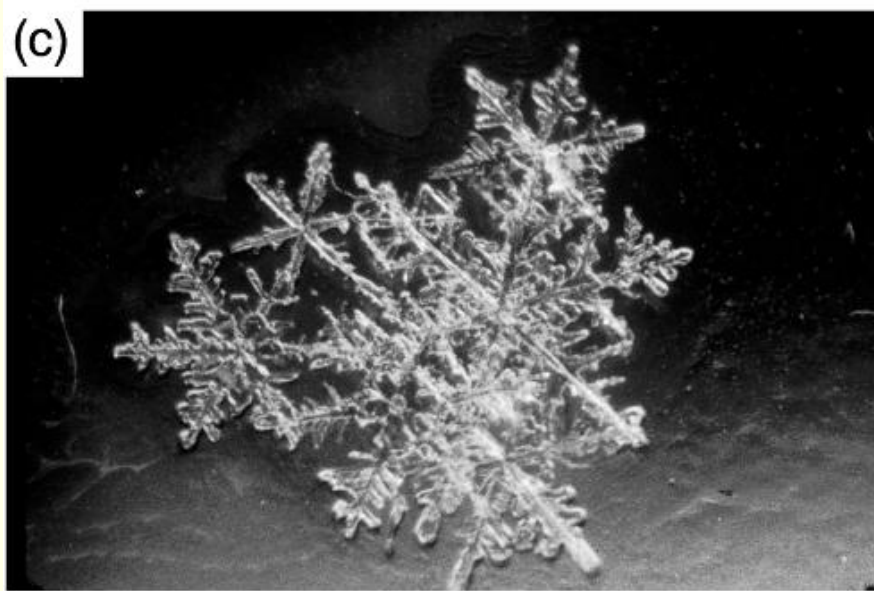
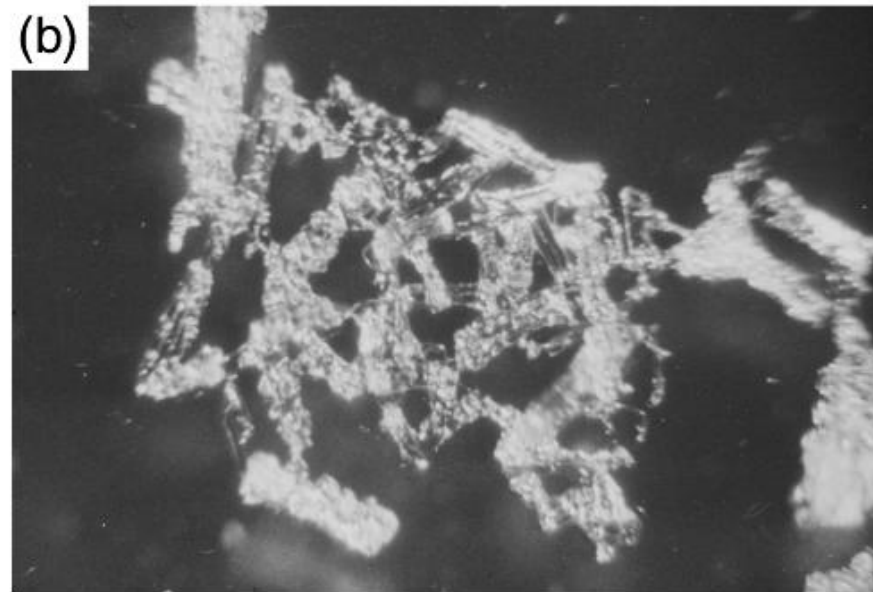
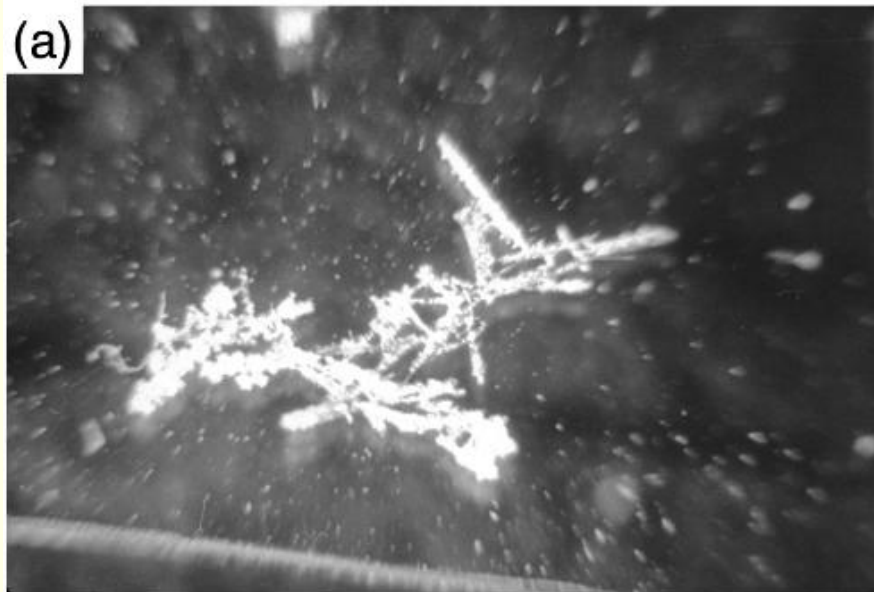
The second factor that influences growth by aggregation is whether or not two ice particles *adhere* when they collide.

The probability of adhesion is determined primarily by:

1. The types of ice particles
2. The temperature

The probability of two colliding crystals adhering increases with increasing temperature as the ice surfaces become more “sticky”.

Some examples of ice particle aggregates are shown below.



Aggregates of (a) rimed needles; (b) rimed columns; (c) dendrites; (d) rimed frozen drops.

Precipitation in Cold Clouds

As early as 1784, Benjamin Franklin suggested that *“much of what is rain, when it arrives at the surface of the Earth, might have been snow, when it began its descent . . . ”*.

Precipitation in Cold Clouds

As early as 1784, Benjamin Franklin suggested that *“much of what is rain, when it arrives at the surface of the Earth, might have been snow, when it began its descent ...”*.

This idea was not developed until the early part of the last century when Alfred Lothar Wegener, in 1911, stated that ice particles would grow preferentially by deposition from the vapour phase in a mixed cloud.

Precipitation in Cold Clouds

As early as 1784, Benjamin Franklin suggested that *“much of what is rain, when it arrives at the surface of the Earth, might have been snow, when it began its descent ...”*.

This idea was not developed until the early part of the last century when Alfred Lothar Wegener, in 1911, stated that ice particles would grow preferentially by deposition from the vapour phase in a mixed cloud.

Subsequently, Tor Bergeron, in 1933, and Theodor Robert Walter Findeisen, in 1938, developed this idea in a more quantitative manner.

Precipitation in Cold Clouds

As early as 1784, Benjamin Franklin suggested that *“much of what is rain, when it arrives at the surface of the Earth, might have been snow, when it began its descent ...”*.

This idea was not developed until the early part of the last century when Alfred Lothar Wegener, in 1911, stated that ice particles would grow preferentially by deposition from the vapour phase in a mixed cloud.

Subsequently, Tor Bergeron, in 1933, and Theodor Robert Walter Findeisen, in 1938, developed this idea in a more quantitative manner.

Since Findeisen carried out his field studies in northwestern Europe, he was led to believe that all rain originates as ice. However, as we have seen, rain can also form in warm clouds by the collision-coalescence mechanism.

Calculations indicate that the growth of ice crystals by deposition of vapour is not sufficiently fast to produce large raindrops.

Calculations indicate that the growth of ice crystals by deposition of vapour is not sufficiently fast to produce large raindrops.

Unlike growth by deposition, the growth rates of an ice particle by *riming* and *aggregation* increase as the ice particle increases in size.

Calculations indicate that the growth of ice crystals by deposition of vapour is not sufficiently fast to produce large raindrops.

Unlike growth by deposition, the growth rates of an ice particle by *riming* and *aggregation* increase as the ice particle increases in size.

A simple calculation shows that a platelike ice crystal, 1 mm in diameter, falling through a cloud with a liquid content of 0.5 g m^{-3} , could develop into a spherical graupel particle about 0.5 mm in radius in a few minutes.

Calculations indicate that the growth of ice crystals by deposition of vapour is not sufficiently fast to produce large raindrops.

Unlike growth by deposition, the growth rates of an ice particle by *riming* and *aggregation* increase as the ice particle increases in size.

A simple calculation shows that a platelike ice crystal, 1 mm in diameter, falling through a cloud with a liquid content of 0.5 g m^{-3} , could develop into a spherical graupel particle about 0.5 mm in radius in a few minutes.

A graupel particle of this size, with a density of 100 kg m^{-3} , has a terminal fall speed of about 1 m s^{-1} and would melt into a drop about $230 \mu\text{m}$ in radius.

Calculations indicate that the growth of ice crystals by deposition of vapour is not sufficiently fast to produce large raindrops.

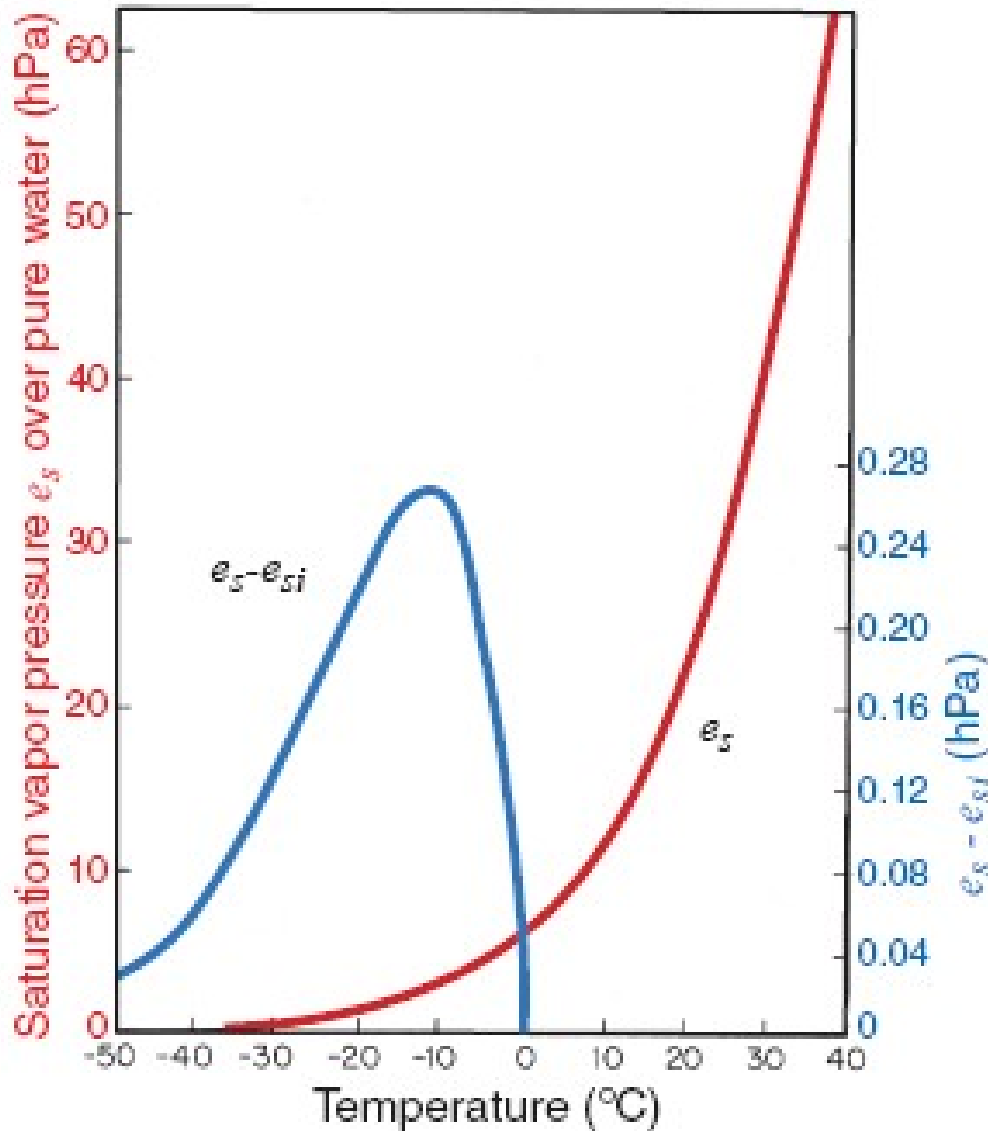
Unlike growth by deposition, the growth rates of an ice particle by *riming* and *aggregation* increase as the ice particle increases in size.

A simple calculation shows that a platelike ice crystal, 1 mm in diameter, falling through a cloud with a liquid content of 0.5 g m^{-3} , could develop into a spherical graupel particle about 0.5 mm in radius in a few minutes.

A graupel particle of this size, with a density of 100 kg m^{-3} , has a terminal fall speed of about 1 m s^{-1} and would melt into a drop about $230 \mu\text{m}$ in radius.

We conclude from these calculations that the growth of ice crystals, first by deposition from the vapour phase in mixed clouds and then by riming and/or aggregation, can produce precipitation-sized particles in about 30 minutes.

Bergeron-Findeisen Process



Saturation vapour pressure over water (Red) and the difference between the saturation pressures over water and over ice (Blue).

Bergeron-Findeisen Process

The vapour pressure over water is greater than over ice
($e_s > e_{si}$).

Bergeron-Findeisen Process

The vapour pressure over water is greater than over ice ($e_s > e_{si}$).

Consider two interconnected chambers, one with super-cooled water, the other with ice (**Draw Picture**).

Bergeron-Findeisen Process

The vapour pressure over water is greater than over ice ($e_s > e_{si}$).

Consider two interconnected chambers, one with super-cooled water, the other with ice (**Draw Picture**).

When the chambers are isolated from each other, there will be fewer molecules of water vapour (per unit volume) over the ice than over the water.

Bergeron-Findeisen Process

The vapour pressure over water is greater than over ice ($e_s > e_{si}$).

Consider two interconnected chambers, one with super-cooled water, the other with ice (**Draw Picture**).

When the chambers are isolated from each other, there will be fewer molecules of water vapour (per unit volume) over the ice than over the water.

If the chambers are then connected, there will be a flow of water vapour from the region over the water to that over the ice.

Bergeron-Findeisen Process

The vapour pressure over water is greater than over ice ($e_s > e_{si}$).

Consider two interconnected chambers, one with super-cooled water, the other with ice (**Draw Picture**).

When the chambers are isolated from each other, there will be fewer molecules of water vapour (per unit volume) over the ice than over the water.

If the chambers are then connected, there will be a flow of water vapour from the region over the water to that over the ice.

Now consider an ice crystal in the atmosphere, surrounded by supercooled water droplets.

Bergeron-Findeisen Process

The vapour pressure over water is greater than over ice ($e_s > e_{si}$).

Consider two interconnected chambers, one with super-cooled water, the other with ice (**Draw Picture**).

When the chambers are isolated from each other, there will be fewer molecules of water vapour (per unit volume) over the ice than over the water.

If the chambers are then connected, there will be a flow of water vapour from the region over the water to that over the ice.

Now consider an ice crystal in the atmosphere, surrounded by supercooled water droplets.

If the air is saturated with respect to water (100% RH), it is supersaturated with respect to ice.

Therefore, water molecules will deposit onto the crystal, thus lowering the relative humidity of the air.

Therefore, water molecules will deposit onto the crystal, thus lowering the relative humidity of the air.

Consequently, water will evaporate from the droplet.

Therefore, water molecules will deposit onto the crystal, thus lowering the relative humidity of the air.

Consequently, water will evaporate from the droplet.

Thus, *ice crystals grow at the expense of water droplets* in a cloud where both are present.

Therefore, water molecules will deposit onto the crystal, thus lowering the relative humidity of the air.

Consequently, water will evaporate from the droplet.

Thus, *ice crystals grow at the expense of water droplets* in a cloud where both are present.

This **Bergeron-Findeisen Process** results in rapid growth of ice crystals and therefore enables precipitation in cold clouds.

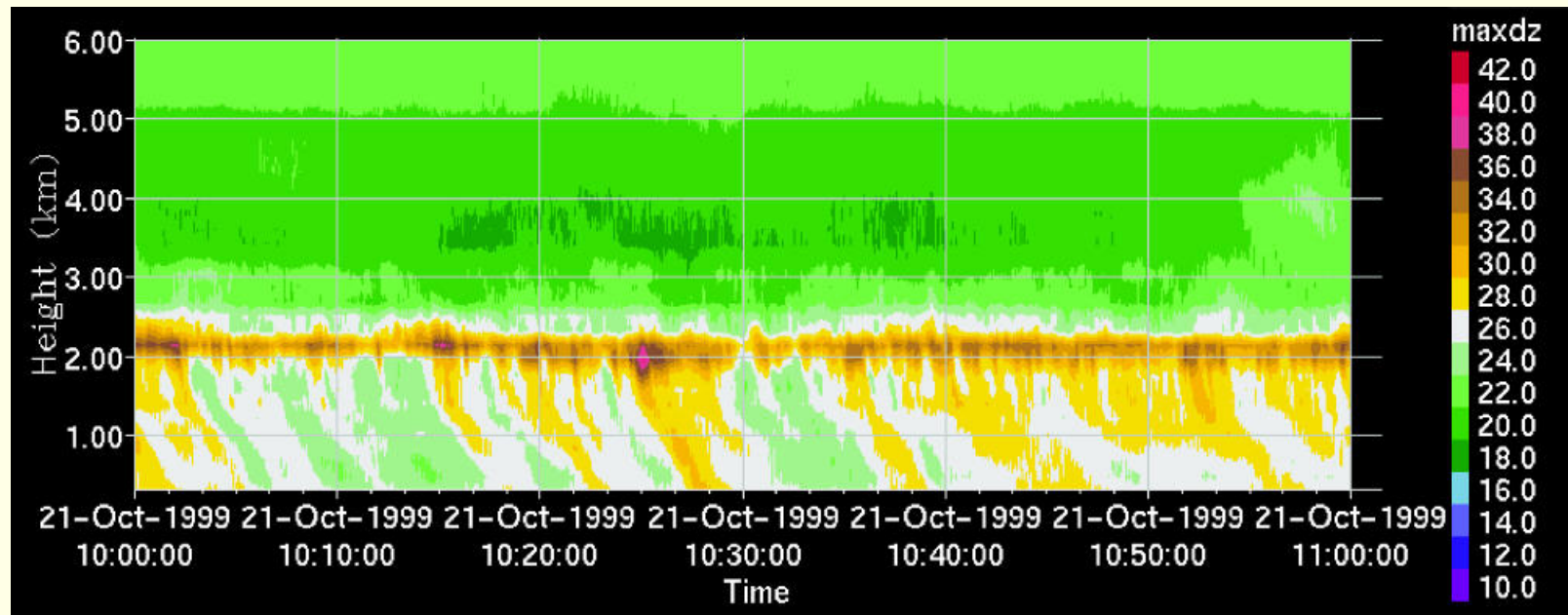
Radar Bright Band

The role of the ice phase in producing precipitation in cold clouds is demonstrated by radar observations.

Radar Bright Band

The role of the ice phase in producing precipitation in cold clouds is demonstrated by radar observations.

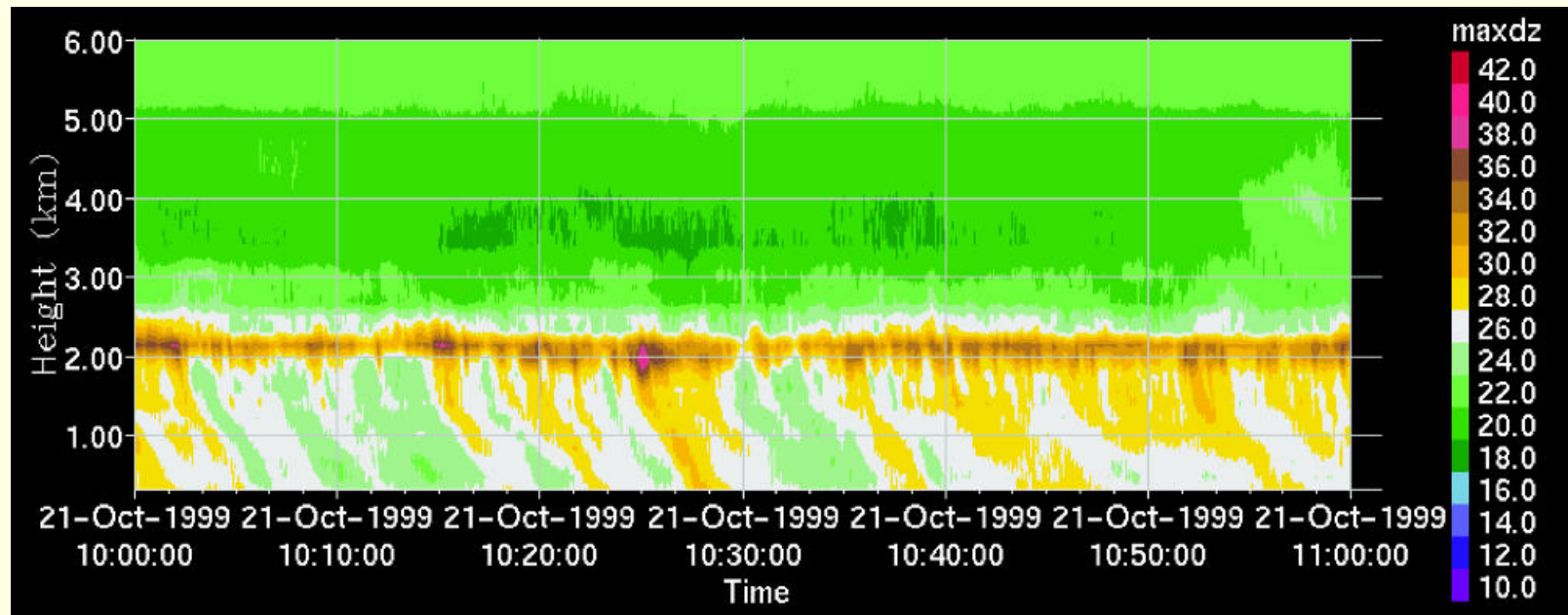
Shown here is a radar image, with the radar antenna pointing vertically upward.



Radar Bright Band

The role of the ice phase in producing precipitation in cold clouds is demonstrated by radar observations.

Shown here is a radar image, with the radar antenna pointing vertically upward.



The horizontal band (brown) just above a height of 2 km was produced by the melting of ice particles. This is referred to as the “**bright band**”.

The curved trails of relatively high reflectivity emanating from the *bright band* are fallstreaks of precipitation, some of which reach the ground.

The curved trails of relatively high reflectivity emanating from the *bright band* are fallstreaks of precipitation, some of which reach the ground.

The radar reflectivity is high around the melting level because, while melting, ice particles become coated with a film of water that greatly increases their radar reflectivity.

The curved trails of relatively high reflectivity emanating from the *bright band* are fallstreaks of precipitation, some of which reach the ground.

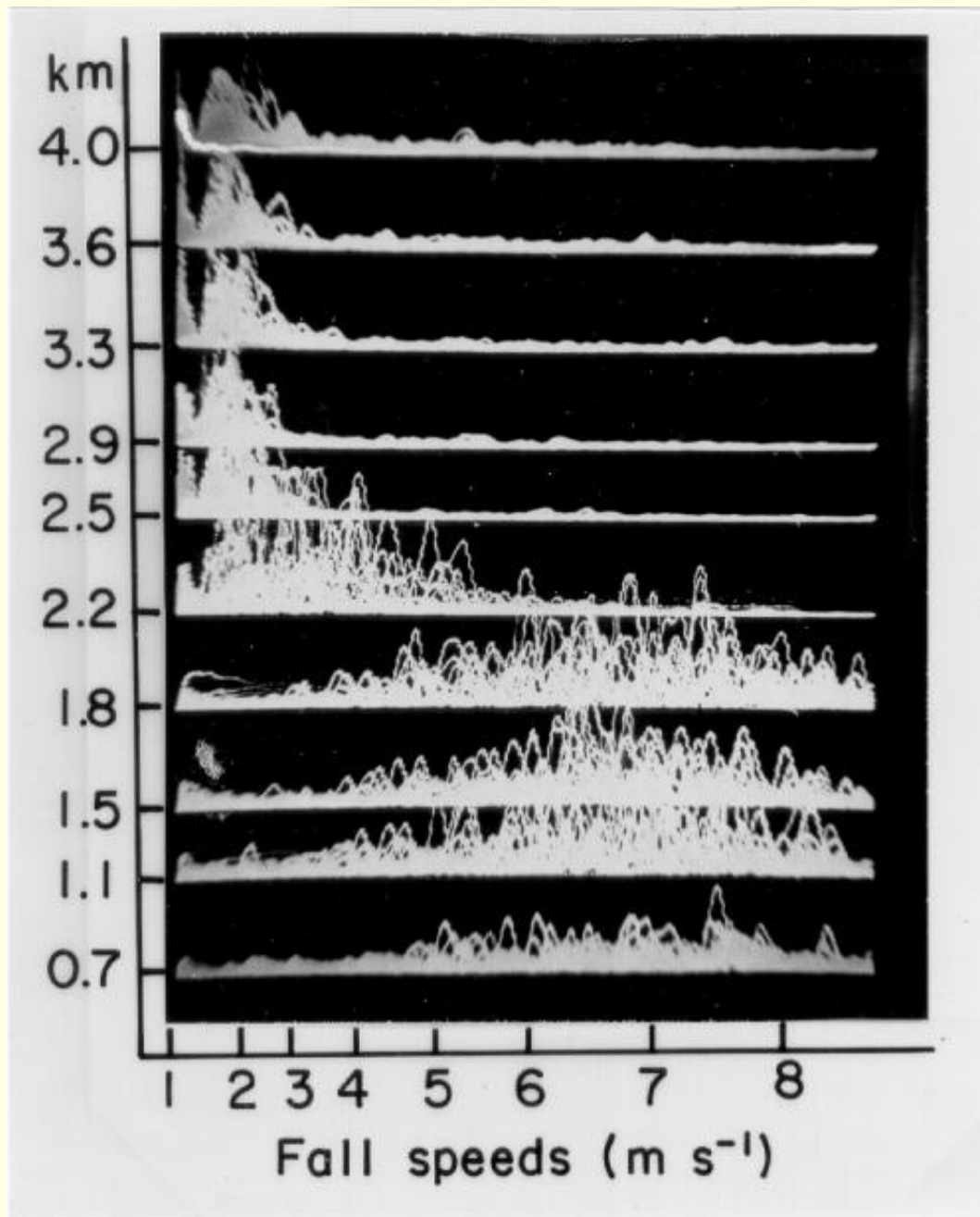
The radar reflectivity is high around the melting level because, while melting, ice particles become coated with a film of water that greatly increases their radar reflectivity.

When the crystals have melted completely they collapse into droplets, and their terminal fall speeds increase so that the concentration of particles is reduced. These changes result in a sharp decrease in radar reflectivity below the melting band.

The sharp increase in particle fall speeds produced by melting is illustrated below, showing the spectrum of fall speeds of precipitation particles measured at various heights with a vertically pointing Doppler radar.

The sharp increase in particle fall speeds produced by melting is illustrated below, showing the spectrum of fall speeds of precipitation particles measured at various heights with a vertically pointing Doppler radar.

At heights above 2.2 km the particles are ice with fall speeds centered around 2 m s^{-1} . At 2.2 km the particles are partially melted, and below 2.2 km there are raindrops with fall speeds centered around 7 m s^{-1} .



Spectra of Doppler fall speeds for precipitation particles at ten heights in the atmosphere. The melting level is at about 2.2 km.

Classification of Solid Precipitation

The growth of ice particles by *deposition*, *riming*, and *aggregation* leads to a very wide variety of solid precipitation particles. A relatively simple classification into ten main classes is given now.

Classification of Solid Precipitation

The growth of ice particles by *deposition*, *riming*, and *aggregation* leads to a very wide variety of solid precipitation particles. A relatively simple classification into ten main classes is given now.

1. *A plate* is a thin, platelike snow crystal the form of which more or less resembles a hexagon or, in rare cases, a triangle. Generally all edges or alternative edges of the plate are similar in pattern and length.

Classification of Solid Precipitation

The growth of ice particles by *deposition*, *riming*, and *aggregation* leads to a very wide variety of solid precipitation particles. A relatively simple classification into ten main classes is given now.

1. *A plate* is a thin, platelike snow crystal the form of which more or less resembles a hexagon or, in rare cases, a triangle. Generally all edges or alternative edges of the plate are similar in pattern and length.
2. *A stellar crystal* is a thin, flat snow crystal in the form of a conventional star. It generally has six arms but stellar crystals with three or twelve arms occur occasionally. The arms may lie in a single plane or in closely spaced parallel planes in which case the arms are interconnected by a very short column.

3. *A column* is a relatively short prismatic crystal, either solid or hollow, with plane, pyramidal, truncated, or hollow ends. Pyramids, which may be regarded as a particular case, and combinations of columns are included in this class.

3. *A column* is a relatively short prismatic crystal, either solid or hollow, with plane, pyramidal, truncated, or hollow ends. Pyramids, which may be regarded as a particular case, and combinations of columns are included in this class.
4. *A needle* is a very slender, needlelike snow particle of approximately cylindrical form. This class includes hollow bundles of parallel needles, which are very common, and combinations of needles arranged in any of a wide variety of fashions.

3. *A column* is a relatively short prismatic crystal, either solid or hollow, with plane, pyramidal, truncated, or hollow ends. Pyramids, which may be regarded as a particular case, and combinations of columns are included in this class.
4. *A needle* is a very slender, needlelike snow particle of approximately cylindrical form. This class includes hollow bundles of parallel needles, which are very common, and combinations of needles arranged in any of a wide variety of fashions.
5. *A spatial dendrite* is a complex snow crystal with fern-like arms which do not lie in a plane or in parallel planes but extend in many directions from a central nucleus. Its general form is roughly spherical.

6. *A capped column* is a column with plates of hexagonal or stellar form at its ends and, in many cases, with additional plates at intermediate positions. The plates are arranged normal to the principal axis of the column. Occasionally only one end of the column is capped in this manner.

6. *A capped column* is a column with plates of hexagonal or stellar form at its ends and, in many cases, with additional plates at intermediate positions. The plates are arranged normal to the principal axis of the column. Occasionally only one end of the column is capped in this manner.
7. *An irregular crystal* is a snow particle made up of a number of small crystals grown together in a random fashion. Generally the component crystals are so small that the crystalline form of the particle can only be seen with the aid of a magnifying glass or microscope.

6. *A capped column* is a column with plates of hexagonal or stellar form at its ends and, in many cases, with additional plates at intermediate positions. The plates are arranged normal to the principal axis of the column. Occasionally only one end of the column is capped in this manner.
7. *An irregular crystal* is a snow particle made up of a number of small crystals grown together in a random fashion. Generally the component crystals are so small that the crystalline form of the particle can only be seen with the aid of a magnifying glass or microscope.
8. *Graupel*, which includes *soft hail*, *small hail*, and *snow pellets*, is a snow crystal or particle coated with a heavy deposit of rime. It may retain some evidence of the outline of the original crystal although the most common type has a form which is approximately spherical.

9. *Ice pellets* (frequently called *sleet* in North America) are transparent spheroids of ice and are usually fairly small. Some ice pellets do not have a frozen center which indicates that, at least in some cases, freezing takes place from the surface inwards.

- 9. *Ice pellets*** (frequently called *sleet* in North America) are transparent spheroids of ice and are usually fairly small. Some ice pellets do not have a frozen center which indicates that, at least in some cases, freezing takes place from the surface inwards.
- 10. *A hailstone*** is a grain of ice, generally having a laminar structure and characterized by its smooth glazed surface and its translucent or milky-white center. Hail is usually associated with those atmospheric conditions which accompany thunderstorms.